

CHEMISTRY-11	Chapter#03 - First Half (3.1 to 3.6) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: oo Min.			

Maximum Marks: oo **(OBJECTIVE TYPE)** Time Allowed: oo Min.

- Which of the followings, will have the same number of molecules at STP?
 - 280 cm³ of CO₂ and 280 cm³ of N₂O
 - 11.2 dm³ of O₂ and 32 g of O₂
 - 44 g of CO₂ and 11.2 dm³ of CO
 - 28 g of N₂ and 5.6 dm³ of oxygen
- The spreading of fragrance of a rose or scent in air is due to:
 - Effusion
 - Diffusion
 - Osmosis
 - Evaporation
- Which of the given relations is the correct one?
 - 1 atm = 1 torr
 - 1 torr = 1 Pa
 - 1 atm = 1 kPa
 - 14.7 psi = 760 mm of Hg
- The graphical representation of Charles' law is called:
 - An isotherm
 - An isobar
 - A negative slope
 - A hyperbolic curve
- How many degrees on Fahrenheit scale are present between the freezing point and the boiling point of water?
 - 100
 - 180
 - 212
 - 380
- The energy demand of one mole of every gas, to increase its temperature by 1 K, is equal to:
 - 0.0821 atm dm³
 - 1.987 cal
 - 8.3143 Nm
 - All
- When a gas is collected over water, its pressure will be:
 - P_{moist} + Aqueous tension
 - P_{dry} + Aqueous tension
 - Aqueous tension - P_{moist}
 - P_{moist} - Aqueous tension
- After every 100 feet depth, the deep sea divers experience _____ greater external pressure on their body.
 - 2 times
 - 3 times
 - 5 times
 - 10 times
- Which gas out of the following will have the slowest rate of diffusion?
 - CH₄
 - O₂
 - Cl₂
 - SO₂

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (12)**
- Give various units of pressure.
 - Throw some light on the factor 1/273 in Charles's law.
 - Calculate the value of R at STP.
 - Define partial pressure.
 - At higher altitude, pilots feel uncomfortable breathing, why?
 - State Graham's law of diffusion. Give its mathematical expression.

SECTION-II

NOTE: Attempt All Questions: (04)

- Q.3: Derive an expression from general gas equation to calculate the density of gas.**