

CHEMISTRY-11	Chapter#06-First Half (6.1.0 - 6.4.3) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The compromise bond distance of two Hydrogen atoms is:
 - 74.4 pm
 - 75.4 pm
 - 71.4 pm
 - 78.4 pm
- The value of third ionization energy of Mg is:
 - 1450 kJmol⁻¹
 - 7730 kJmol⁻¹
 - 7850 kJmol⁻¹
 - 1890 kJmol⁻¹
- CsF has ionic character:
 - 60%
 - 80%
 - 92%
 - 100%
- VSEPR theory was proposed by:
 - Kossel
 - Nylholm & Gillespie
 - Lewis
 - Sidgwick
- The shape of SnCl₂ molecule is:
 - Linear
 - Angular
 - Trigonal planar
 - Tetrahedral
- From left to right in a period, the atomic radii generally:
 - Increase
 - Decrease
 - Remain same
 - All
- The covalent radius of H is:
 - 37 pm
 - 37.7 pm
 - 74 pm
 - 75.4 pm
- The electron pair repulsions decrease in the order:
 - B.P. – B.P. > B.P. – L.P. > L.P. – L.P.
 - L.P. – L.P. > B.P. – B.P. > B.P. – L.P.
 - B.P. – L.P. > B.P. – B.P. > L.P. – L.P.
 - L.P. – L.P. > L.P. – B.P. > B.P. – B.P.
- The F-N-F bond angle in NF₃ is:
 - 109.5°
 - 107.5°
 - 104.5°
 - 102°

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

- Q.2: Give brief answers to the following questions:** (12)
- Define bond distance?
 - Why cationic radius is smaller than the parent atom?
 - Why does I.E. decrease from top to bottom in a group?
 - Differentiate between polar and non-polar bond.
 - Write down two postulates of VSEPR theory.
 - Why NH₃ molecule and NH₄⁺ ions have different structures?

SECTION-II

NOTE: Attempt All Questions:

(04)

- Q.3: Define coordinate covalent bond. Explain with examples.**