

CHEMISTRY-11	Chapter#04-First Half (4.1-4.3) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Acetone and chloroform are soluble into each other due to:
 - Intermolecular hydrogen bonding
 - Ion-dipole interaction
 - Instantaneous dipole
 - All of the above
- When water freezes at 0°C, its density decreases due to:
 - Cubic structure of ice
 - Empty spaces present in the structure of ice
 - Change of bond lengths
 - Change of bond angles
- Vapor pressure of a substance does not depend upon:
 - Temperature
 - Physical state of matter
 - Intermolecular forces
 - Surface area
- Which pair will have ion-dipole forces?
 - Water + NaCl
 - Water + C₂H₆
 - Water + CO₂
 - Water + Acetic acid
- 88.6°C is the boiling point of:
 - Methane
 - Ethane
 - Hexane
 - Decane
- Which one is the strongest acid?
 - HF
 - HCl
 - HBr
 - HI
- The hydrides of which group have the lowest boiling points?
 - VII A
 - IV A
 - V A
 - VIA
- The number of H-bonds between adenine and thymine in DNA is:
 - 1
 - 2
 - 3
 - 4
- What is the correct order with respect to ΔH_v values?
 - F₂ > Cl₂ > Br₂ > I₂
 - Br₂ > I₂ > Cl₂ > F₂
 - I₂ > Br₂ > Cl₂ > F₂
 - Cl₂ > F₂ > Br₂ > I₂

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (12)

- What are instantaneous dipole-induced dipole forces or London dispersion forces?
- H₂O is a liquid while H₂S is a gas at room temperature. Explain.
- Why are ethyl alcohol and carboxylic acids soluble in water but hydrocarbons are not?
- Evaporation occurs at all temperatures. Explain.
- What is vacuum distillation? What are its advantages?
- Heat of sublimation of I₂ is very high. Explain with reason.

SECTION-II

NOTE: Attempt All Questions:

(04)

03: What are liquid crystals. Give their uses.