CHEMISTRY-11	Chapter#06-First Half (6.1.0 - 6.4.3) Test-1		
	Name:	Class:	ID:
Date: / /	**Marks Total: 25**	**Marks Obtained:**	
Time Allowed: 40 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

1) The octet rule is not followed by:

(a) NH_3 (b) Cl_2 (c) CCl_4 (d) PCl_5

2) First ionization energy of Mg atom is:

(a) +738 $kJmol^{-1}$ (b) +1450 $kJmol^{-1}$ (c) -349 $kJmol^{-1}$ (d) -500 $kJmol^{-1}$

3) The highest electronegative element in the periodic table is:

(a) Oxygen (b) Nitrogen (c) Chlorine (d) Fluorine

4) Which of the following has linear structure?

(a) CO_2 (b) NH_3 (c) CH_4 (d) H_2O

5) Which one is the smallest of all?

(a) Atomic radius (b) Covalent radius (c) Cationic radius (d) Anionic radius

6) Which one is correct order with respect to ionic radius?

(a) $Na^+ < Mg^{2+} < Al^{3+}$ (b) $Mg^{2+} < Al^{3+} < Na^+$ (c) $Al^{3+} < Mg^{2+} < Na^+$ (d) $Mg^{2+} < Na^+ < Al^{3+}$

7) The elements of which group show abnormal trend of ionization energy?

(a) II A (b) III A (c) VI A (d) Both 'b' & 'c'

8) The difference between the E.N. values of bonded atoms is an index to:

(a) Polarity of bond (b) Stability

(c) Metallic character (d) Conductivity

9) In NH_4^+, each bond is _____ coordinate.

(a) 25% (b) 33% (c) 50% (d) 100%

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions: **(12)**

i. Define chemical bond. Give the cause of chemical bond formation.

ii. Why the radius of an atom cannot be determined precisely?

iii. Why anionic size is greater than the parent atom?

iv. Define electron affinity. Give an example.

v. Define electronegativity. How does it vary along the period and down the group?

vi. Define co-ordinate covalent bond with the help of two examples.

SECTION-II

NOTE: Attempt All Questions: **(04)**

Q.3: Discuss the main postulates of VSEPR theory and explain the structures of NH_3 on the basis of this theory.