

CHEMISTRY-11	Chapter#04-First Half (4.1-4.3) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 40 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- London dispersion forces are the only forces present among the:
 - Molecules of water in liquid state.
 - Atoms of helium in gaseous state at high temperature.
 - Molecules of solid iodine.
 - Molecules of hydrogen chloride gas.
- In order to mention the boiling point of water at 110°C, the external pressure should be:
 - Between 760 torr and 1200 torr
 - Between 200 torr and 760 torr
 - 765 torr
 - Any value of pressure
- Density of H₂O is maximum at:
 - 0°C
 - 2°C
 - 1°C
 - 4°C
- The boiling point of glycerin at 1 atmospheric pressure is:
 - 210°C
 - 270°C
 - 290°C
 - 300°C
- Which one has dipole-dipole forces?
 - H₂O
 - HCl
 - C₆H₆
 - All
- Which halogen is a solid at room temperature?
 - F₂
 - O₂
 - Br₂
 - I₂
- The strongest H-bond is present between:
 - H and F
 - H and N
 - H and O
 - All equal
- The double helix of DNA molecules has a diameter of:
 - 10-12 Å
 - 18-20 Å
 - 20-22 Å
 - 25-27 Å
- Liquid crystals were discovered by:
 - Peter Debye
 - William Crooks
 - Fritz London
 - Fredrick Reinitzer

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(12)

- Differentiate between intermolecular and intramolecular forces.
- Define polarizability.
- Why is HF weaker than all the halogen acids?
- Why does gasoline evaporate faster than water?
- Define molar heat of sublimation with one example.
- How do liquid crystals act as temperature sensors?

SECTION-II

NOTE: Attempt All Questions:

(04)

03: Define hydrogen bonding? Explain the structure of ice with the help of hydrogen bonding.