

CHEMISTRY-11	Chapter#06 (Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Which of the following molecules has zero dipole moment?
 (a) NH_3 (b) CHCl_3 (c) H_2O (d) BF_3
- The highest electronegative element in the periodic table is:
 (a) Oxygen (b) Nitrogen (c) Chlorine (d) Fluorine
- Which one of the followings has the highest bond order?
 (a) O_2^{+1} (b) O_2^{+2} (c) O_2^{-1} (d) O_2^{-2}
- According to Lewis theory, ionic bond is formed by the complete transfer of electron from an atom with _____ I.E. to another atom with _____ E.A.
 (a) Low, low (b) High, high (c) High, low (d) Low, high
- The molecular geometry of SO_2 resembles to that of:
 (a) H_2O (b) CO_2 (c) SO_3 (d) NH_3
- The S.I. unit of dipole moment is:
 (a) Debye (b) Magnetron (c) Nm (d) Cm

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

- Q.2: Give brief answers to the following questions:** **(16)**
- Bond distance is the compromise distance between two atoms. Explain.
 - Why the size of Cl atom is smaller than Cl^- ion?
 - Define co-ordinate covalent bond with the help of two examples.
 - Why does lone pair of electrons occupy more space than bond pairs?
 - What is π -bond? Give an example.
 - Differentiate between atomic and molecular orbitals.
 - Why is O_2 paramagnetic?
 - Isomerism is shown by covalent compounds, not by ionic compounds. Why?

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.3:** What is ionization energy. Discuss variation of ionization energy in periodic table. Also explain how ionization energy is an index to metallic character?
- Q.4:** Write a note on sp hybridization and give example of ethyne.