

CHEMISTRY-11	Chapter#05 (Complete) Test-6		
	Name:	Class:	ID:
Date: / /	Marks Total: 40	Marks Obtained:	
Time Allowed: 75 Min.			

Maximum Marks: 08

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- In the ground state of an atom, the electron is present:
 - In the nucleus
 - In the second shell
 - Nearest to the nucleus
 - Farthest from the nucleus
- Cathode rays can be generated at the pressure of:
 - 1 torr
 - 0.1 torr
 - 0.01 torr
 - 0.001 torr
- The number of neutrons present in ${}^{39}_{19}\text{K}$ is:
 - 39
 - 18
 - 20
 - 19
- Maximum number of electrons in an orbital is:
 - 6
 - 10
 - 14
 - 2
- The cathode rays are negatively charged. This was established by _____ in 1897.
 - William Crooks
 - J. Thomson
 - J. Perrin
 - R. A. Millikan
- Slow neutrons move with energy less than:
 - 1 eV
 - 2 eV
 - 3 eV
 - 4 eV
- Which is the correct order?
 - $r_2 - r_1 > r_3 - r_2 > r_4 - r_3 > \dots$
 - $r_4 - r_3 < r_3 - r_2 < r_2 - r_1$
 - $r_2 - r_1 < r_3 - r_2 < r_4 - r_3 < \dots$
 - None
- How many total electrons are in N atom, which have 'm' value +1?
 - 2
 - 3
 - 5
 - 1

Maximum Marks: 32

(SUBJECTIVE TYPE)

Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(20)

- Cathode rays have reducing effect. Explain.
- Write four properties of positive rays.
- Calculate the mass of an electron when its e/m is $1.758 \times 10^{11} \text{ C kg}^{-1}$.
- Define frequency, wavelength and wave number.
- What is atomic emission spectrum?
- What are the defects in Bohr's model?
- Give importance of Mosley's law.
- What is principal quantum number? Give its significance.
- Why are p-orbitals called triply fold degenerate orbitals?
- What are degenerate orbitals?

SECTION-II

NOTE: Attempt All Questions:

(12)

- Write J. J. Thomson's method to measure the e/m value of electron.
- Write a note on Heisenberg's uncertainty principle.
- Draw the shapes of p and d orbitals.