

CHEMISTRY-11	Chapter#04 (Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 40	Marks Obtained:	
Time Allowed: 75 Min.			

Maximum Marks: 08 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Amorphous solids:**
 - Have sharp melting points.
 - Undergo clean cleavage when cut with knife.
 - Have perfect arrangement of atoms.
 - Can possess small regions of orderly arrangement of atoms.
- Dipole-dipole forces are present among:**
 - Molecules of Iodine
 - Atoms of Neon in gaseous state
 - Chloroform molecules
 - CCl_4 molecules
- When water freezes, its volume increases:**
 - 10%
 - 9%
 - 15%
 - 18%
- Cholesteryl benzoate turns into milky liquid at:**
 - 144°C
 - 145°C
 - 146°C
 - 147°C
- Dipole-dipole forces are _____ effective as covalent bond.**
 - 1%
 - 5%
 - 20%
 - 50%
- Water has minimum volume at:**
 - 0°C
 - 4°C
 - 25°C
 - 100°C
- The molar heat of vaporization of water is:**
 - 40.6 kJ mol^{-1}
 - 574 kJ mol^{-1}
 - $40.6 \text{ kcal mol}^{-1}$
 - $574 \text{ kcal mol}^{-1}$
- MgO is isomorphic to:**
 - NaF
 - NaNO_3
 - ZnSO_4
 - K_2CrO_4

Maximum Marks: 32 **(SUBJECTIVE TYPE)** Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (20)

- What are instantaneous dipole-induced dipole forces or London dispersion forces?
- Why is the boiling point of H_2O greater than that of HF, although F is more electronegative than O atom?
- Why do we feel cooling after bathing?
- What is vacuum distillation? What are its advantages?
- Define molar heat of fusion with one example.
- Give four uses of liquid crystals.
- Define anisotropy with example.
- What is habit of crystal? Can habit be changed?
- Transition temperature is lower than melting point, why?

x. What are the types of covalent solids?

SECTION-II

NOTE: Attempt All Questions:

(12)

03: Define hydrogen bonding? Explain the structure of ice with the help of hydrogen bonding.

04: What are liquid crystals. Give their uses.

05: Explain covalent solids with their properties.