

CHEMISTRY-11	Chapter#04 (Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- NH₃ shows maximum boiling point among the hydrides of Vth group elements due to:**
 - Very small size of nitrogen
 - Lone pair of electrons present on nitrogen
 - Enhanced electronegative character of nitrogen
 - Pyramidal structure of NH₃
- Which of the followings is a pseudo solid?**
 - CaF₂
 - Glass
 - NaCl
 - All
- Down the VII-A group, polarizability generally:**
 - Decreases
 - Increases
 - Remains same
 - Unpredictable
- When $a=b \neq c$ and $\alpha = \beta = 90^\circ, \gamma = 120^\circ$ then system is:**
 - Cubic
 - Triclinic
 - Hexagonal
 - Monoclinic
- The hydrides of which group have the highest boiling points?**
 - IV A
 - V A
 - VI A
 - VII A
- In solid state, the I-I bond length is:**
 - 271.5 pm
 - 266.6 pm
 - 154 pm
 - None

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (16)

- What are Debye's forces or dipole-induced dipole forces?
- Why fluorine is a gas while iodine is a solid at room temperature?
- In a very cold winter, fish in garden ponds owe their lives to hydrogen bonding. Explain.
- Why do earthenware vessels keep water cool?
- Define with example dynamic equilibrium.
- Cleavage is anisotropic property, explain.
- Define transition temperature with an example.
- Iodine dissolves readily in tetrachloromethane. Why?

SECTION-II

NOTE: Attempt All Questions:

(08)

03: What is hydrogen bonding? Discuss hydrogen bonding in biological compounds.

04: What are ionic solids? Give their properties.