

<b>CHEMISTRY-11</b>	<b>Chapter#04 (Complete) Test-2</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 40</b>	<b>Marks Obtained:</b>	
Time Allowed: 75 Min.			

Maximum Marks: 08 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- When water freezes at  $0^{\circ}\text{C}$ , its density decreases due to:
  - Cubic structure of ice
  - Empty spaces present in the structure of ice
  - Change of bond lengths
  - Change of bond angles
- Diamond is a bad conductor because:
  - It has a tight structure.
  - It has a high density.
  - There are no free electrons present in the crystal of diamond to conduct electricity.
  - Is transparent to light.
- The boiling point of water at Murree hills:
  - $90^{\circ}\text{C}$
  - $98^{\circ}\text{C}$
  - $100^{\circ}\text{C}$
  - $120^{\circ}\text{C}$
- Which pair will have Debye's forces?
  - Water + NaCl
  - Water +  $\text{CH}_3\text{OH}$
  - Water +  $\text{CO}_2$
  - Water + Acetic acid
- Hydrogen bonding is \_\_\_\_\_ times less than that of covalent bond.
  - 10
  - 20
  - 30
  - 5
- The number of amino acids in each turn of the helix is:
  - 10
  - 18
  - 27
  - 40
- Which one is not amorphous solid?
  - Glass
  - Plastic
  - Glue
  - Sugar
- The unit cell angle present between a and c is called:
  - $\alpha$
  - $\beta$
  - $\gamma$
  - $\Phi$

Maximum Marks: 32 **(SUBJECTIVE TYPE)** Time Allowed: 65 Min.

### SECTION-I

- Q.2: Give brief answers to the following questions: (20)**
- What is the role of London dispersion forces in the liquefaction of noble gases?
  - $\text{H}_2\text{O}$  is a liquid while  $\text{H}_2\text{S}$  is a gas at room temperature. Explain.
  - Why does ice float on the surface of water?
  - Define evaporation. Name the factors that control the rate of evaporation?
  - Why does water boil at different temperatures at Murree hills and on the Mount Everest?
  - Why is  $\Delta H_s$  of a substance greater than  $\Delta H_v$ ?
  - Differentiate between polymorphism and allotropy.

- viii. 58.5 amu is the formula mass of NaCl, not its molecular mass. Explain.
- ix. Molecular solids are soft and easily compressible. Justify.
- x. Why does the electrical conductivity of metal decrease with increase in temperature?

## **SECTION-II**

**NOTE:** Attempt All Questions:

(12)

- 03: What is the effect of external pressure on the boiling point of a substance? Give example.
- 04: Differentiate between isomorphism and polymorphism with suitable examples.
- 05: Write four properties of metallic crystals.