

<b>CHEMISTRY-11</b>	<b>Chapter#04 (Complete) Test-2</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 30</b>	<b>Marks Obtained:</b>	
Time Allowed: 60 Min.			

Maximum Marks: 06

**(OBJECTIVE TYPE)**

Time Allowed: 10 Min.

**NOTE:** Tick The Correct Option:

- In order to mention the boiling point of water at  $110^{\circ}\text{C}$ , the external pressure should be:
  - Between 760 torr and 1200 torr
  - Between 200 torr and 760 torr
  - 765 torr
  - Any value of pressure
- The molecules of  $\text{CO}_2$  in dry ice form the:
  - Ionic crystals
  - Covalent crystals
  - Molecular crystals
  - Any type of crystals
- The boiling point of glycerin at 1 atmospheric pressure is:
  - $210^{\circ}\text{C}$
  - $270^{\circ}\text{C}$
  - $290^{\circ}\text{C}$
  - $300^{\circ}\text{C}$
- The carbon atom in diamond is \_\_\_\_\_ hybridized.
  - $sp^3$
  - $sp^2$
  - $sp$
  - $dsp^2$
- Which halogen is a liquid at room temperature?
  - $\text{Cl}_2$
  - $\text{Br}_2$
  - $\text{F}_2$
  - None
- The transition temperature of Sulphur ( $\text{S}_8$ ) is:
  - $13.2^{\circ}\text{C}$
  - $95.5^{\circ}\text{C}$
  - $128^{\circ}\text{C}$
  - $32.02^{\circ}\text{C}$

Maximum Marks: 24

**(SUBJECTIVE TYPE)**

Time Allowed: 50 Min.

**SECTION-I**

**Q.2: Give brief answers to the following questions: (16)**

- What are dipole-dipole forces?
- Why is HF weaker than all the halogen acids?
- Why do we feel cool when we sit under fan after bathing?
- Heat of sublimation of  $\text{I}_2$  is very high. Explain with reason.
- Amorphous solids like glass are also called super cooled liquids, explain.
- Ionic crystals are highly brittle. Justify.
- Diamond is a bad conductor of electricity while graphite is a good conductor of electricity, why?
- Freshly cut metals give lustrous surface. Explain.

**SECTION-II**

**NOTE:** Attempt All Questions:

(08)

**03:** Define vapour pressure. Describe manometric method for the measurement of vapor pressure of a liquid.

**04:** What are molecular solids? Give their properties and examples.