

CHEMISTRY-11	Chapter#04 (Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 40	Marks Obtained:	
Time Allowed: 75 Min.			

Maximum Marks: 08

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- London dispersion forces are the only forces present among the:
 - Molecules of water in liquid state.
 - Atoms of helium in gaseous state at high temperature.
 - Molecules of solid iodine.
 - Molecules of hydrogen chloride gas.
- Dipole-dipole forces are present among:

(a) Molecules of Iodine	(b) Atoms of Neon in gaseous state
(c) Chloroform molecules	(d) CCl_4 molecules
- Density of H_2O is maximum at:

(a) $0^\circ C$	(b) $2^\circ C$	(c) $-1^\circ C$	(d) $4^\circ C$
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- Transition temperature of KNO_3 is:

(a) $13.2^\circ C$	(b) $95.5^\circ C$	(c) $128^\circ C$	(d) $32.2^\circ C$
(a) 10	(b) 20	(c) 30	(d) 5
- The coiling of the amino acids chain in proteins is maintained by hydrogen bonding between:

(a) N and H	(b) O and H	(c) F and H	(d) Cl and H
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- The molar heat of vaporization of water is:

(a) 40.6 kJ mol^{-1}	(b) 574 kJ mol^{-1}	(c) $40.6 \text{ kcal mol}^{-1}$	(d) $574 \text{ kcal mol}^{-1}$
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- Which one is anisotropic with respect to electrical conductivity?

(a) Diamond	(b) Graphite	(c) NaCl	(d) Cu metal
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- The crystals of LiF are:

(a) Primitive cubes	(b) Body centered cubes
(c) Face centered cubes	(d) Hexagonal

Maximum Marks: 32

(SUBJECTIVE TYPE)

Time Allowed: 65 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(20)

- Why are intermolecular forces weaker than intramolecular forces?
- Why do the boiling points of Noble gases increase down the group?
- Why are ethyl alcohol and carboxylic acids soluble in water but hydrocarbons are not?
- Why do we feel cooling after bathing?
- Why does water boil at different temperatures at Murree hills and on the Mount Everest?
- Give four uses of liquid crystals.
- Define symmetry. Name any two symmetry elements.
- Transition temperature is lower than melting point, why?
- Why NaCl and CsCl have different structures?
- Why does the electrical conductivity of metal decrease with increase in temperature?

SECTION-II

NOTE: Attempt All Questions:

(12)

06: What is the effect of external pressure on the boiling point of a substance? Give example.

09: Define the following with examples:

i) Amorphous solids ii) Habit of crystals iii) Allotropy iv) Transition temperature

12: Explain covalent solids with their properties.