

CHEMISTRY-11	Chapter#04 (Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 60 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Acetone and chloroform are soluble into each other due to:
 - Intermolecular hydrogen bonding
 - Ion-dipole interaction
 - Instantaneous dipole
 - All of the above
- Ionic solids are characterized by:
 - Low melting points
 - High vapour pressures
 - Good conductivity in solid state
 - Solubility in polar solvents
- Water may boil at 120°C when external pressure is:
 - 369 torr
 - 700 torr
 - 760 torr
 - 1489 torr
- The structure of NO_3^- ion is:
 - Square planar
 - Triangular
 - Tetrahedral
 - Pentagonal
- The number of H-bonds between adenine and thymine in DNA is:
 - 1
 - 2
 - 3
 - 4
- Atomic solid is another name of:
 - Ionic solids
 - Molecular solids
 - Covalent solids
 - Metallic solids

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 50 Min.

SECTION-I

Q.2: Give brief answers to the following questions: (16)

- Differentiate between intermolecular and intramolecular forces.
- Define polarizability.
- Evaporation causes cooling. Explain.
- Why is food cooked earlier in the pressure cooker?
- What are liquid crystals?
- Differentiate between polymorphism and isomorphism.
- Define lattice energy. Give example.
- Sodium is softer than copper but both are very good electrical conductors. Justify.

SECTION-II

NOTE: Attempt All Questions:

(08)

- Define and explain London forces. Describe the factors affecting the London dispersion forces.
- Explain the structure of diamond.