

CHEMISTRY-11	Chapter#03 (Complete) Test-3D		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The order of the rate of diffusion of gasses NH_3 , SO_2 , Cl_2 , and CO_2 is:
 - $\text{NH}_3 > \text{SO}_2 > \text{Cl}_2 > \text{CO}_2$
 - $\text{NH}_3 > \text{CO}_2 > \text{SO}_2 > \text{Cl}_2$
 - $\text{Cl}_2 > \text{SO}_2 > \text{CO}_2 > \text{NH}_3$
 - $\text{NH}_3 > \text{CO}_2 > \text{Cl}_2 > \text{SO}_2$
- The deviation of gas form ideal behavior is maximum at:
 - -10°C and 5.0 atm
 - -10°C and 2.0 atm
 - 100°C and 2.0 atm
 - 0°C and 2.0 atm
- The commonly used unit of pressure by meteorologists is:
 - Atmosphere
 - Pascal
 - Millibar
 - psi
- In Fahrenheit scale, absolute zero is equal to:
 - -459.69°F
 - -241.16°F
 - -549.288°F
 - -308.16°F
- Plasma was discovered by:
 - Van der Waals
 - J.J. Thomson
 - Clausius
 - William Crooks
- Which one is an example of natural plasma?
 - Auroras
 - Fluorescent bulbs
 - Neon sings
 - None

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

- Q.2: Give brief answers to the following questions: (16)**
- What do you mean by the term isotherm?
 - Describe different scales of thermometry.
 - How does Dalton's law explain the process of respiration?
 - Describe law of distribution of velocities.
 - Derive Avogadro's law from KMT.
 - Define compressibility factor.
 - What are the units of 'a' and 'b'?
 - What are the characteristics of plasma?

SECTION-II

NOTE: Attempt All Questions:

(08)

- Q.0: Calculate the mass of 1 dm^3 of NH_3 gas at 30°C and 1000 mm Hg pressure, considering that NH_3 is behaving ideally.**
- Q.0: a) Two moles of NH_3 are enclosed in a 5 dm^3 flask at 27°C . Calculate pressure exerted by the gas assuming that:**
- It behaves like an ideal gas
 $a = 4.17 \text{ atm dm}^6 \text{ mol}^{-2}$
 - It behaves like a real gas
 $b = 0.0371 \text{ dm}^3 \text{ mol}^{-1}$