

CHEMISTRY-11	Chapter#03 (Complete) Test-1C		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06

(OBJECTIVE TYPE)

Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Pressure remaining constant at which temperature the volume of a gas will become twice of what it is at 0°C:
 - 546°C
 - 200°C
 - 546 K
 - 273 K
- The molar volume of CO₂ is maximum at:
 - STP
 - 127°C and 1 atm
 - 0°C and 2 atm
 - 273°C and 2 atm
- The SI unit of pressure is expressed in:
 - Nm⁻¹
 - Nm⁻²
 - Nm⁻³
 - mm Hg
- The volume of every gas will become theoretically zero at:
 - 0°C
 - 273 K
 - 273 K
 - 273°C
- The property of the gas which is inversely proportional to temperature:
 - Pressure
 - Volume
 - Density
 - Molar mass
- The compressibility factor is:
 - $\frac{PV}{RT}$
 - $\frac{PV}{T}$
 - $\frac{PM}{dRT}$
 - $\frac{PV}{nRT}$

Maximum Marks: 24

(SUBJECTIVE TYPE)

Time Allowed: 40 Min.

SECTION-I

Q.2: Give brief answers to the following questions:

(16)

- Why are liquids less common than other forms of matter?
- Throw some light on the factor 1/273 in Charles's law.
- Prove that $d = PM/RT$.
- State Avogadro's law.
- State Graham's law of diffusion. Give its mathematical expression.
- Define and explain Joule Thomson effect.
- Water vapours don't behave ideally at 273 K, explain.
- Give important uses of plasma.

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: What is Joule-Thomson effect? Explain Linde's method for the liquefaction of gases.

Q.4: What pressure is exerted by a mixture of 2.00g of H₂ and 8.00g of N₂ at 273K in a 10 dm³ vessel?