CHEMISTRY-11	Chapter#02 (Complete) Test-2 (SMART Syllabus)		
	Name:	Class:	ID:
Date: / /	**Marks Total: 20**	**Marks Obtained:**	
Time Allowed: 30 Min.			

Maximum Marks: 10 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

i. **Solvent extraction method is a partially useful technique for separation when the product to be separated is:**

(a) Non-volatile or thermally unstable (b) Volatile or thermally stable

(c) Non-volatile or thermally stable (d) Volatile or thermally unstable

ii. **The substance which does not show the process of sublimation?**

(a) $K_2Cr_2O_7$ (b) Iodine (c) Naphthalene (d) NH_4Cl

iii. **I_2, in CCl_4 solvent, shows color:**

(a) Blue (b) Brown (c) Purple (d) Pink

iv. **During the process of sublimation, the substance under experiment is heated on:**

(a) Water bath (b) Sand bath (c) Hot finger (d) Hot air

v. **The Greek word 'khromatos' means:**

(a) Color writing (b) Color making (c) Color mixing (d) Both 'a' & 'b'

vi. **In adsorption chromatography, the stationary phase is a:**

(a) Solid (b) Liquid (c) Gas (d) Both 'a' & 'b'

vii. **The chromatographic operation is stopped when the solvent front has risen upto _____ of the paper.**

(a) 1/4 (b) 2/3 (c) 1/3 (d) 3/4

viii. **The greater the R_f value of a component, the _____ is the distance travelled by it from the original spot.**

(a) Greater (b) Smaller (c) Both (d) None

Maximum Marks: 12 **(SUBJECTIVE TYPE)** Time Allowed: 20 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (12)

i. How can naphthalene be purified?

ii. What is ether extraction?

iii. State distribution law. Explain with an example.

iv. Define chromatography. Give formula of distribution co-efficient.

v. Differentiate between partition chromatography and adsorption chromatography.

vi. Define chromatogram.