

<b>CHEMISTRY-11</b>	<b>Chapter#01 (Complete) Test-3</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 25</b>	<b>Marks Obtained:</b>	
Time Allowed: 45 Min.			

Maximum Marks: 09

**(OBJECTIVE TYPE)**

Time Allowed: 15 Min.

**NOTE:** Tick The Correct Option:

- The number of moles of  $\text{CO}_2$  which contain 8.0 g of Oxygen:  
(a) 0.25 (b) 0.50 (c) 1.0 (d) 1.50
- The largest number of molecules are present in:  
(a) 3.6 g of  $\text{H}_2\text{O}$  (b) 4.8 g of  $\text{C}_2\text{H}_5\text{OH}$  (c) 2.8 g of  $\text{CO}$  (d) 5.4 g of  $\text{N}_2\text{O}_5$
- 1 gram atom of carbon is equal to:  
(a) 1 gram C (b) 12 gram C (c) 12.0000 amu C (d)  $6.02 \times 10^{23}$  gram C
- One gram molecule of  $\text{H}_2\text{SO}_4$  dissolves completely in water to give:  
(a) One gram ion (b) Two gram ions (c) Three gram ions (d) 98 gram ions
- A chemical equation does not tell about \_\_\_\_\_ of reaction.  
(a) Rate (b) Conditions (c) Efficiency (d) All
- A limiting reactant is one which:  
(a) Is present in less amount in grams. (b) Is consumed earlier in the reaction.  
(c) Decreases the rate of the reaction. (d) Gives maximum product.
- A closed container is filled with air. The nitrogen and oxygen of the air are made to react to form NO, according to the equation  $\text{N}_2 + \text{O}_2 \rightarrow 2\text{NO}$ . What will be the limiting reactant?  
(a)  $\text{N}_2$  (b)  $\text{O}_2$  (c) Both (d) None
- In rusting of iron, the limiting reactant is:  
(a)  $\text{O}_2$  (b) Moisture (c) Iron (d) Sunlight
- The amount of product obtained in a chemical reaction is called:  
(a) Actual yield (b) Theoretical yield (c) Percentage yield (d) All

Maximum Marks: 16

**(SUBJECTIVE TYPE)**

Time Allowed: 30 Min.

**SECTION-I**

**Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS:**

**(12)**

- Justify that  $\text{N}_2$  and  $\text{CO}$  have same number of electrons, protons and neutrons.
- 180 g of glucose and 342 g of sucrose have the same number of molecules but different number of atoms present in them. Give reason.
- Calculate the mass in grams of  $10^{-3}$  moles of water.
- Define Avogadro's number. Give an equation to relate the Avogadro's number and the mass of an element.
- What are the limitations of a chemical equation?
- Limiting reactant controls the amount of the product. Explain?

**SECTION-II**

**NOTE:** Attempt All Questions:

**(04)**

**Q.3: What is the difference between actual yield and theoretical yield? Why actual yield is less than theoretical yield?**