

CHEMISTRY-11	Chapter#01 (Complete) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 45 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 15 Min.

NOTE: Tick The Correct Option:

- The mass of one mole of electron is:
 (a) 1.008 mg (b) 0.55 mg (c) 0.184 mg (d) 1.673 mg
- One mole of SO_2 contains:
 (a) 6.02×10^{23} atoms of Oxygen (b) 18.1×10^{23} molecules of SO_2
 (c) 6.02×10^{23} atoms of Sulphur (d) 4 gram atoms of SO_2
- The number of carbon atoms in 22.0 g of CO_2 is:
 (a) 3.01×10^{23} (b) 6.02×10^{23} (c) 3.01×10^{22} (d) 6.02×10^{22}
- The mass of one atom of hydrogen is:
 (a) 1.008 g (b) 1.67×10^{-22} g (c) 1.008 amu (d) Both 'b' & 'c'
- The volume occupied by 1.6 g of O_2 at STP is:
 (a) 22.4 dm³ (b) 11.2 dm³ (c) 2.24 dm³ (d) 1.12 dm³
- Stoichiometric calculations are not true for:
 (a) Irreversible reactions (b) Reversible reactions
 (c) Spontaneous reactions (d) Non-spontaneous reactions
- If 4 g of H_2 is allowed to react with 64 g of O_2 , the amount of H_2O formed will be:
 (a) 32 g (b) 64 g (c) 18 g (d) 36 g
- The amount of product obtained in a chemical reaction is called:
 (a) Actual yield (b) Theoretical yield (c) Percentage yield (d) All
- Actual yield is mostly less than the theoretical yield because of:
 (a) Inexperienced worker (b) The wastage of some product
 (c) Competing side reaction (d) All
- Which type of yield tells about the efficiency of the reaction?
 (a) Theoretical (b) Actual (c) Percentage (d) All

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (12)

- Define gram ion.
- One mg of K_2CrO_4 has thrice the number of ions than the number of formula units when ionized in water.
- Calculate the number of gram atoms in 0.1 kg of silicon.
- One mole H_2SO_4 requires 2 moles of NaOH to neutralize it. Explain.
- Define Stoichiometry. Give two assumptions for stoichiometric calculations.
- Why in some reactions, one of the reactants is used deliberately in excess quantity?

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: An unknown metal M reacts with S to form a compound with a formula M_2S_3 . If 3.12 g of M reacts with exactly 2.88 g of sulphur, what are the names of metal M and the compound M_2S_3 ?