CHEMISTRY-11	Chapter#01 (Complete) Test-1		
	Name:	Class:	ID:
Date: / /	**Marks Total: 25**	**Marks Obtained:**	
Time Allowed: 45 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 15 Min.

NOTE: Tick The Correct Option:

1. **27 g of Al will react completely with how much mass of O_2 to produce Al_2O_3?**
   (a) 8 g of Oxygen (b) 16 g of Oxygen (c) 32 g of Oxygen (d) 24 g of Oxygen
2. **The volume occupied by 1.4 g of N_2 at S.T.P is:**
   (a) 2.24 dm^3 (b) 22.4 dm^3 (c) 1.12 dm^3 (d) 112 cm^3
3. **The mass of two moles of electrons is:**
   (a) 1.10 mg (b) 1.00 mg (c) 0.184 mg (d) 1.673 mg
4. **The number of gram ions produced by the complete dissociation of one gram molecule of H_3PO_4 is:**
   (a) One (b) Two (c) Three (d) Four
5. **Which gas will occupy greater volume at STP?**
   (a) 10 g H_2 (b) 50 g CH_4 (c) 100 g O_2 (d) 200 g of SO_2
6. **Stoichiometric calculations obey:**
   (a) Law of conservation of energy (b) Law of conservation of mass
   (c) Law of definite proportions (d) Both 'b' & 'c'
7. **If there is an expensive reactant, it should act as:**
   (a) Limiting reactant (b) Excess reactant (c) Catalyst (d) None
8. **In burning of wood, the limiting reactant is:**
   (a) O_2 (b) N_2 (c) Air (d) Wood
9. **A practically inexperienced worker may affect:**
   (a) Theoretical yield (b) Actual yield (c) Percentage yield (d) Both 'b' & 'c'

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (12)

i. Define gram molecule.
ii. Mg atom is twice heavier as an atom of carbon. Explain.
iii. Calculate the number of gram atoms (moles) in 0.1 g of sodium.
iv. When H_2SO_4 is dissolved in water, the number of ions is unequal, but the number of charges is equal. Explain.
v. Define limiting reactant. Give an example.
vi. Why in most reactions, actual yield remains less than the theoretical yield?

SECTION-II

NOTE: Attempt All Questions: (04)

Q.3: A sample of 0.600 mole of metal M reacts completely with excess of fluorine to form 46.8 g of MF_2.

(a) How many moles of F are present in the sample of MF_2 that forms?

(b) Which element is rpresented by the symbol M?