

CHEMISTRY-11	Chapter#01 (Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 45 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 15 Min.

NOTE: Tick The Correct Option:

- Many elements have fractional atomic masses. This is because:
 - The mass of the atom is itself fractional.
 - Atomic masses are average masses of isobars.
 - Atomic masses are average masses of isotopes.
 - Atomic masses are average masses of isotopes proportional to their relative abundance.
- Cadmium has isotopes:
 - 5
 - 6
 - 9
 - 2
- In combustion analysis, H_2O vapours are absorbed by:
 - $Mg(ClO_4)_2$
 - $Mg(ClO_3)_3$
 - $Mg(ClO_3)_2$
 - 50% KOH
- Which one is the molecule of the element?
 - S_8
 - HCl
 - CO_2
 - NH_3
- Each hemoglobin molecule is _____ times heavier than hydrogen atom.
 - 1000
 - 10000
 - 68000
 - 100000
- The molecular ions can be generated by passing _____ through a gas.
 - High energy electron beams
 - Alpha particles
 - X-rays
 - All
- The elements having _____ almost possess more than two stable isotopes:
 - Odd atomic number
 - Even atomic number
 - Odd mass number
 - Even mass number
- Combustion analysis is an example of:
 - Qualitative analysis
 - Quantitative analysis
 - Salt analysis
 - None
- Which one of the following compounds has the highest percentage of oxygen by weight?
 - CH_3OH
 - C_2H_5OH
 - HCOOH
 - H_2O

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS:

(12)

- Why can atom not be seen with optical microscope?
- What are molecular ions? How are they generated?
- Why do isotopes have same chemical but different physical properties?
- Write down any four methods used for the separation of isotopes.
- What is the justification of two equally strong peaks in the mass spectrum for bromine, while for iodine only one peak at 127amu is indicated?
- A compound may have same empirical and molecular formulas. Justify giving examples.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: What is mass spectrometer? How is it used to determine the relative atomic masses of isotopes?