

<b>CHEMISTRY-11</b>	<b>Chapter#01 (Complete) Test-2</b>		
	Name:	Class:	ID:
Date: / /	<b>Marks Total: 25</b>	<b>Marks Obtained:</b>	
Time Allowed: 45 Min.			

Maximum Marks: 09 **(OBJECTIVE TYPE)** Time Allowed: 15 Min.

**NOTE:** Tick The Correct Option:

- Which of the following statements is not true?
  - Isotopes with even atomic masses are comparatively abundant.
  - Isotopes with odd atomic masses are comparatively abundant.
  - Isotopes with even atomic masses and even atomic numbers are comparatively abundant.
  - Isotopes with even atomic masses and odd atomic numbers are comparatively abundant.
- The atomicity of  $\text{NH}_3$  is:
  - 1
  - 2
  - 3
  - 4
- Ascorbic acid is vitamin:
  - A
  - B
  - C
  - E
- J. Berzelius determined the atomic \_\_\_\_\_ of the elements.
  - Number
  - Radii
  - Masses
  - Structure
- Masses of atoms range from:
  - $10^{-27}$  g to  $10^{-25}$  g
  - $10^{-30}$  kg to  $10^{-27}$  kg
  - $10^{-27}$  kg to  $10^{-25}$  kg
  - $10^{-30}$  g to  $10^{-27}$  g
- The formation of a positive ion is \_\_\_\_\_ process.
  - An endothermic
  - An exothermic
  - A heterothermic
  - None
- Which one is a tri-negative ion?
  - Sulphate
  - Permanganate
  - Phosphate
  - Dichromate
- Isotopes have different number of:
  - Electrons
  - Protons
  - Neutrons
  - None
- The potential difference applied between perforated accelerating plates in Dempster's mass spectrometer is:
  - 500 - 1000 V
  - 500 - 2000 V
  - 1000 - 2000 V
  - 1500 - 2000 V

Maximum Marks: 16 **(SUBJECTIVE TYPE)** Time Allowed: 30 Min.

### SECTION-I

**Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (12)**

- Differentiate between an 'atom' and a 'molecule'.
- The formation of a positive ion is an endothermic reaction. Explain.
- What is amu?
- What is the function of magnetic field in mass spectrometer?
- Calculate the percentage of nitrogen in  $\text{NH}_2\text{CONH}_2$ .
- What are the steps to determine the empirical formula?

### SECTION-II

**NOTE:** Attempt All Questions: (04)

**Q.3:** A sample of liquid consisting of carbon, hydrogen and oxygen was subjected to combustion analysis. 0.5439 g of the compound gave 1.039 g of  $\text{CO}_2$ , 0.6369 g of  $\text{H}_2\text{O}$ . Determine the empirical formula of the compound?