

CHEMISTRY-11	Chapter#01 (Complete) Test-1		
	Name:	Class:	ID:
Date: / /	Marks Total: 25	Marks Obtained:	
Time Allowed: 45 Min.			

Maximum Marks: 09

(OBJECTIVE TYPE)

Time Allowed: 15 Min.

NOTE: Tick The Correct Option:

- Isotopes differ in:**
 - Properties which depend upon mass
 - Arrangement of electrons in orbitals
 - Chemical properties
 - The extent to which they may be affected in electromagnetic field.
- One atomic mass unit equals:**
 - 1.661×10^{-27} kg
 - 1.60×10^{-19} kg
 - 9.1×10^{-31} kg
 - 1.661×10^{-30} kg
- Nickel has isotopes:**
 - 3
 - 5
 - 6
 - 11
- The empirical formula of glucose is:**
 - CHO
 - $C_2H_4O_2$
 - CH_2O
 - $C_6H_{12}O_6$
- A full stop may contain about _____ atoms.**
 - One million
 - Two Million
 - Five million
 - 15 million
- On C-12 scale, the relative atomic mass of ^{12}C is:**
 - 12 amu
 - 12.00 amu
 - 12.000 amu
 - 12.0000 amu
- Which set of elements constitute about 50 % of earth crust?**
 - $^{12}C, ^{16}O, ^{24}Mg, ^{40}Ca, ^{56}Fe$
 - $^{16}O, ^{24}Mg, ^{32}S, ^{40}Ca, ^{56}Fe$
 - $^{16}O, ^{24}Mg, ^{28}Si, ^{32}S, ^{40}Ca$
 - $^{16}O, ^{24}Mg, ^{28}Si, ^{40}Ca, ^{56}Fe$
- Which organic compound can be analyzed by combustion analysis?**
 - C_2H_5OH
 - C_2H_5Cl
 - $C_2H_5NH_2$
 - All
- Which compound has same empirical and molecular formulas?**
 - H_2O
 - NH_3
 - C_2H_5OH
 - All

Maximum Marks: 16

(SUBJECTIVE TYPE)

Time Allowed: 30 Min.

SECTION-I

Q.2: GIVE BRIEF ANSWERS TO THE FOLLOWING QUESTIONS: (12)

- What are macromolecules? Give an example.
- Define isotopes and isotopy.
- Define mass spectrum.
- Neon has atomic mass 20.18 amu. But no individual atom in Neon has this value of mass. Explain.
- Differentiate between qualitative and quantitative analyses.
- Differentiate between empirical formula and molecular formula.

SECTION-II

NOTE: Attempt All Questions:

(04)

Q.3: The combustion analysis of an organic compound shows it to contain 65.44% carbon, 5.50% hydrogen and 29.06% oxygen. What is the empirical formula of the compound? If the molecular mass of this compound is $110.15 \text{ g mol}^{-1}$. Calculate the molecular formula of the compound.