

CHEMISTRY-11	Chapter#01 (Complete) Test-4		
	Name:	Class:	ID:
Date: / /	Marks Total:	30	Marks Obtained:
Time Allowed: 50 Min.			

Maximum Marks: 06 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- Many elements have fractional atomic masses. This is because:
 - The mass of the atom is itself fractional.
 - Atomic masses are average masses of isobars.
 - Atomic masses are average masses of isotopes.
 - Atomic masses are average masses of isotopes proportional to their relative abundance.
- Which is not a molecular ion:
 - He⁺
 - CH₄⁺
 - NH₄⁺
 - CO⁺
- Nickel has isotopes:
 - 3
 - 5
 - 6
 - 11
- J. Berzelius determined the atomic _____ of the elements.
 - Number
 - Radii
 - Masses
 - Structure
- Each hemoglobin molecule is made up of nearly:
 - 1000 atoms
 - 10,000 atoms
 - 68,000 atoms
 - 100,000 atoms
- Ascorbic acid is also called:
 - Vitamin A
 - Vitamin B
 - Vitamin C
 - Vitamin D

Maximum Marks: 24 **(SUBJECTIVE TYPE)** Time Allowed: 40 Min.

SECTION-I

Q.2: Give Brief Answers To The Following Short Questions: (16)

- What are molecular ions? How are they generated?
- Define isotopes and isotopy.
- What is the function of ionization chamber in mass spectrometer?
- Why can the percentage of oxygen not be calculated directly in combustion analysis?
- Mg atom is twice heavier as an atom of carbon. Explain.
- Calculate the number of gram atoms (moles) in 0.1 g of sodium.
- Law of conservation of mass has to be obeyed during stoichiometric calculations. Explain.
- Limiting reactant controls the amount of the product. Explain?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: Describe combustion analysis to determine the percentage of C, H and O in the organic compound.

Q.3: What is the difference between actual yield and theoretical yield? Why actual yield is less than theoretical yield?