

CHEMISTRY-11	Chapter#01 (Complete) Test-3		
	Name:	Class:	ID:
Date: / /	Marks Total:	30	Marks Obtained:
Time Allowed: 50 Min.			

Maximum Marks: 06 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- 27 g of Al will react completely with how much mass of O₂ to produce Al₂O₃?
 (a) 8 g of Oxygen (b) 16 g of Oxygen (c) 32 g of Oxygen (d) 24 g of Oxygen
- A limiting reactant is the one which:
 (a) Is taken in lesser quantity in grams as compared to other reactants.
 (b) Is taken in lesser quantity in volume as compared to the other reactants.
 (c) Gives the maximum amount of the product which is required.
 (d) Gives the minimum amount of the product under consideration.
- One atomic mass unit equals:
 (a) 1.661×10^{-27} kg (b) 1.60×10^{-19} kg (c) 9.1×10^{-31} kg (d) 1.661×10^{-30} kg
- In combustion analysis, H₂O vapours are absorbed by:
 (a) Mg(ClO₄)₂ (b) Mg(ClO₃)₃ (c) Mg(ClO₃)₂ (d) 50% KOH
- Aluminium forms _____ positive ion.
 (a) Mono (b) Di (c) Tri (d) Tetra
- Which isotope of Neon has the smallest peak in mass spectrum?
 (a) ²⁰Ne (b) ²¹Ne (c) ²²Ne (d) ²⁴Ne

Minimum Marks: 24 **(SUBJECTIVE TYPE)** Time Allowed: 40 Min.

SECTION-I

Q.2: Give Brief Answers To The Following Short Questions: (12)

- What are macromolecules? Give an example.
- What are mono-isotopic elements?
- Differentiate between empirical formula and molecular formula.
- Define gram atom, giving examples.
- Calculate the number of gram atoms in 0.1 kg of silicon.
- Define molar volume.
- Why in some reactions, one of the reactants is used deliberately in excess quantity?
- Differentiate between actual yield and theoretical yield.

SECTION-II

NOTE: Attempt All Questions: (08)

- Q.3: What is mass spectrometer? How is it used to determine the relative atomic masses of isotopes?
- Q.4: Write down various steps to calculate the empirical formula of a compound.