

CHEMISTRY-11	Chapter#01 (Complete) Test-2		
	Name:	Class:	ID:
Date: / /	Marks Total: 30	Marks Obtained:	
Time Allowed: 50 Min.			

Maximum Marks: 06 **(OBJECTIVE TYPE)** Time Allowed: 10 Min.

NOTE: Tick The Correct Option:

- The mass of one mole of electron is:
 - 1.008 mg
 - 0.55 mg
 - 0.184 mg
 - 1.673 mg
- The number of moles of CO_2 which contain 8.0 g of Oxygen:
 - 0.25
 - 0.50
 - 1.0
 - 1.50
- One mole of SO_2 contains:
 - 6.02×10^{23} atoms of Oxygen
 - 18.1×10^{23} molecules of SO_2
 - 6.02×10^{23} atoms of Sulphur
 - 4 gram atoms of SO_2
- The percentage of Nitrogen in ammonia is:
 - 82.35 %
 - 46.67 %
 - 92%
 - 78%
- Ascorbic acid is vitamin:
 - A
 - B
 - C
 - E
- In rusting of iron, the limiting reactant is:
 - O_2
 - Moisture
 - Iron
 - Sunlight

Maximum Marks: 24 **(SUBJECTIVE TYPE)** Time Allowed: 40 Min.

SECTION-I

Q.2: Give Brief Answers To The Following Short Questions:

(12)

- What is amu?
- Define mass spectrum.
- Neon has atomic mass 20.18 amu. But no individual atom in Neon has this value of mass. Explain.
- What are the steps to determine the empirical formula?
- 58.5 amu is the formula mass of NaCl, but not the molecular mass. Why?
- One mole H_2SO_4 requires 2 moles of NaOH to neutralize it. Explain.
- Define limiting reactant. Give an example.
- How can the efficiency of chemical reaction be calculated? Or What is the percentage yield?

SECTION-II

NOTE: Attempt All Questions:

(08)

Q.3: A well-known ideal gas is enclosed in a container having volume 500 cm^3 at S.T.P. Its mass comes out to be 0.72 g. What is the molar mass of this gas?

Q.4: When lime stone ($CaCO_3$) is roasted, quick lime (CaO) is produced according to the following equation. The actual yield of CaO is 2.5 kg, when 4.5 kg of lime stone is roasted. What is the percentage yield of this reaction? $CaCO_3(g) \longrightarrow CaO(g) + CO_2(g)$